

Thermal Properties Module

Outline

- Heat of Fusion
  
- Heat Capacity
  
- Thermal Expansion
  
- Thermal Conductivity
  
- Thermal Diffusivity

Heat Capacity

- Heat Per Degree
  - Can Be Based on Mass, Moles, Volume.

$$C_P = \frac{dQ}{dT}$$

Titanium (Ti)  
is Basis for  
This Example

$$T_M = 1660^\circ\text{C}$$

$$50\%T_M = 700^\circ\text{C}$$

$$C_P = 2.4 \frac{J}{\text{cm}^3 K}$$

$$\Delta T = 675^\circ\text{C}$$

$$Q = VC_P\Delta T$$

$$= (121\text{cm}^3) \left( 2.4 \frac{J}{\text{cm}^3 K} \right) (675\text{C})$$

$$= 1.6 \times 10^3 J$$

$$RT = 25^\circ\text{C}$$

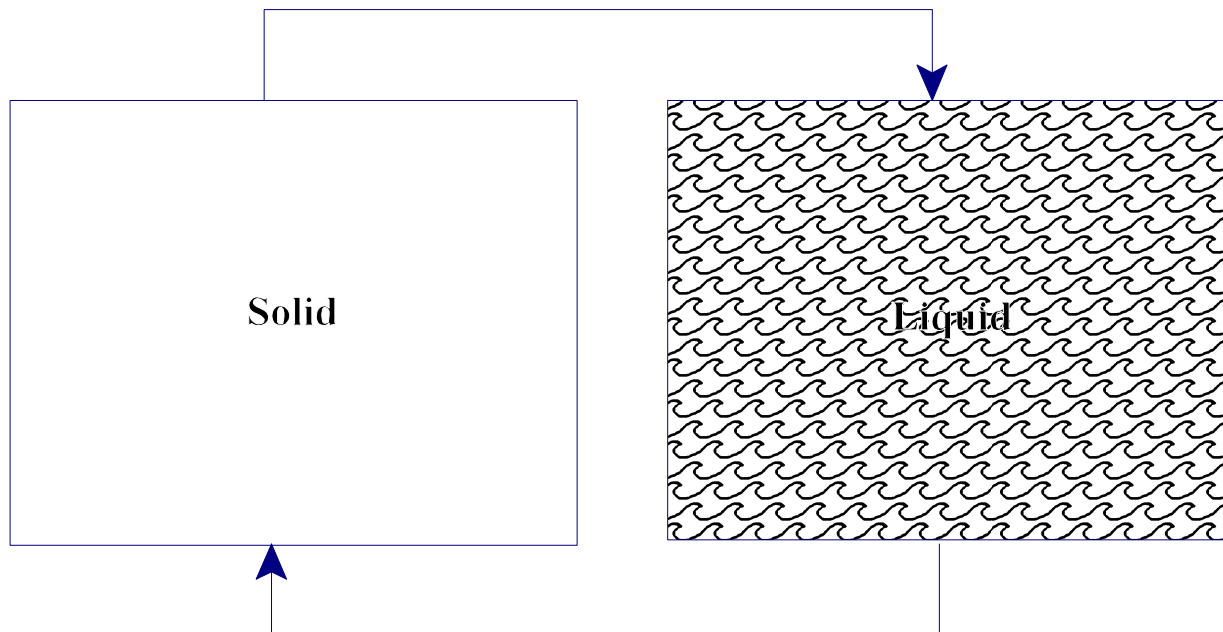
$$D = 3.2\text{cm}, L = 15\text{cm}, V = 121\text{cm}^3$$

Heat of Fusion  $\Delta H_F$ 

- First Law of Thermodynamics
  - Heat at Constant Pressure Required to Melt a Given Amount of Material at  $T_M$
  - Heat at Constant Pressure Expelled When a Given Amount of Solidifies at  $T_M$

$$\begin{aligned}
 T_M &= 660^\circ\text{C} & Q &= V\Delta H_F \\
 V &= 1953 \text{ cm}^3 & &= (1953 \text{ cm}^3) \left( 1080 \frac{\text{J}}{\text{cm}^3} \right) \\
 \Delta H_F &= 1080 \frac{\text{J}}{\text{cm}^3} & &= 2.1 \times 10^6 \text{ J}
 \end{aligned}$$

**$2.1 \times 10^6 \text{ J}$  Required**



**$2.1 \times 10^6 \text{ J}$  Given Off**

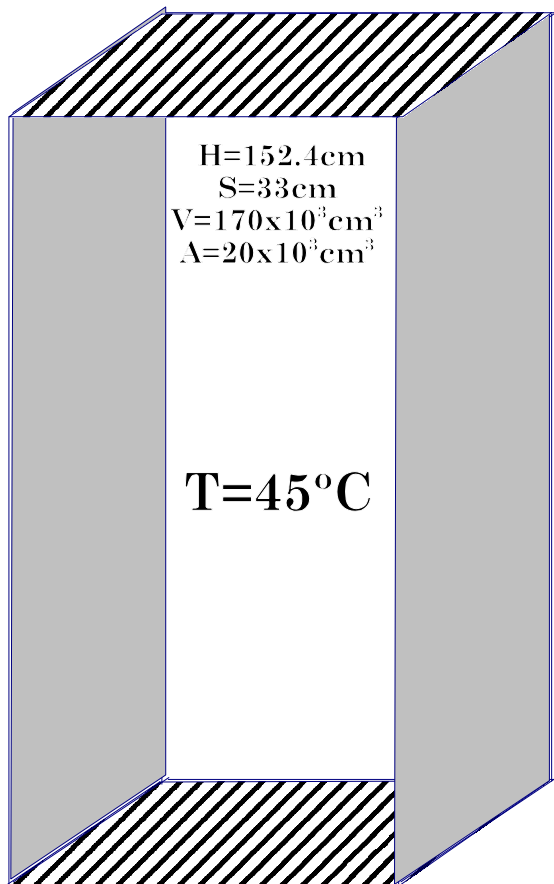
**Aluminum (Al)  
is Basis for  
This Example**

Thermal Conductivity

- Thermal Conductivity

$$\frac{Power}{Area} = -k \frac{dT}{dx}$$

Copper (Cu) is  
Basis for This  
Example



45°C

$$\frac{dT}{dx} = \frac{30\text{C}}{3.5\text{cm}} = 8.6 \frac{\text{C}}{\text{cm}}$$

$$k = 4 \frac{\text{W}}{\text{cmK}}$$

$$\frac{Power}{Area} = -k \frac{dT}{dx}$$

$$\frac{P}{A} = 34.4 \frac{\text{W}}{\text{cm}^2}$$

$$P_{Loss} = 688\text{kW}$$

3.5cm

15°C

Thermal Diffusivity

- Definition

$$\alpha = \frac{k}{C}$$

- Unit Analysis

$$\alpha = \frac{k}{C} = \frac{\frac{J}{(s)(cm)(K)}}{\frac{J}{(cm^3)(K)}} = \frac{cm^2}{s}$$

- Implications

$$\frac{\partial T}{\partial t} = \alpha \left( \frac{\partial^2 T}{\partial x^2} + \frac{\partial^2 T}{\partial y^2} + \frac{\partial^2 T}{\partial z^2} \right)$$

Thermal Expansion

- Thermal Expansion or Contraction

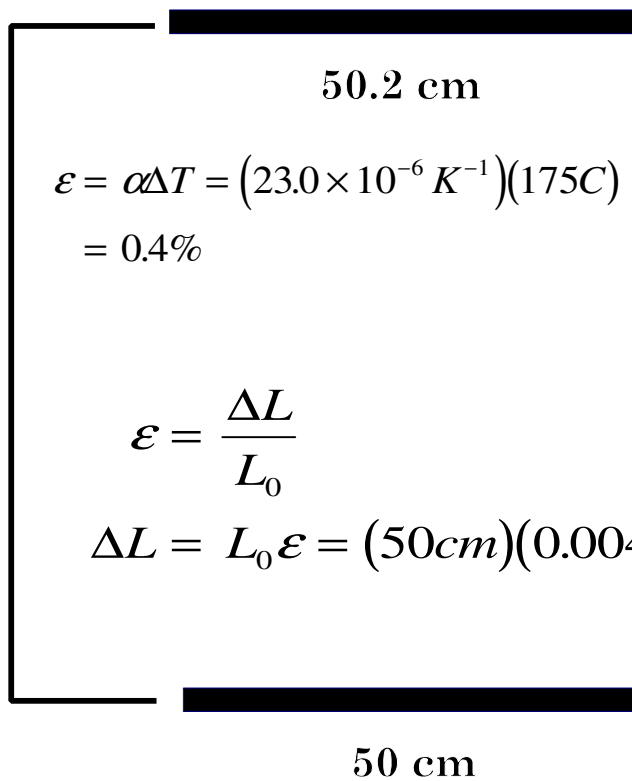
$$\epsilon = \frac{\Delta L}{L_0} = \alpha \Delta T$$

$$94\% \quad T_M = 200^\circ\text{C}$$

$$T_M = 232^\circ\text{C}$$

$$\alpha = 23.0 \times 10^{-6} \text{ K}^{-1}$$

$$\text{RT} = 25^\circ\text{C}$$



Tin (Sn) is Basis for This Example

Summary and Wrap Up

- Heat of Fusion
  
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